Name: _____ Date: _____

Unit 5 - Topic 1

Chemical Reactions

Identify the type of reaction shown in each of the following chemical equations. Your choices are *synthesis, decomposition, single replacement (SR), double replacement (DR), or combustion.*

1	$C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$	
2	$AgNO_3(aq) + KCl(aq) \rightarrow AgCl(s) + KNO_3(aq)$	
3	$2H_2(g) + O_2(g) \rightarrow 2H_2O(g)$	
4	$2AgNO_3(aq) + Cu(s) \rightarrow Cu(NO_3)_2(aq) + 2Ag(s)$	
5	$CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$	
6	$NaI(aq) + Cl_2(g) \rightarrow 2NaCl(aq) + I_2(g)$	
7	$2H_2O_2 \rightarrow 2H_2O + O_2$	
8	$AgNO_3 + NaCl \rightarrow AgCl + NaNO_3$	
9	$2KBr + Cl_2 \rightarrow 2KCl + Br_2$	
10	$2HgO \rightarrow 2Hg + O_2$	

Write a BALANCED chemical equation AND identify what type of reaction is taking place.

11	zinc + lead (II) nitrate yields zinc nitrate + lead
12	aluminum bromide + calcium chloride yields aluminum chloride + calcium bromide
13	potassium chlorate when heated yields potassium chloride + oxygen gas