

Name: _____

Date: _____

Unit 1 - Topic 1

Calculating Errors of Measurements

Complete the following problems using the ESA Method (Equation, Substitute with units, Answer with units). You can find the % error equation on your chemistry reference tables (Table T).

If you are trying to get to New York city and you are 5 cm off course, you will still arrive in New York City. If you are trying to take some dirt out of your eye and you are 5 cm off course, you are working on the wrong eye. The size of the error is the same. The size of the error compared to the size of the target is not the same. The actual size of the error – the difference between the observed value and the true value – is known as the absolute error. The sign of the absolute error is not important. The size of the error is more important than whether the value is over or under.

- Observed value - value based on laboratory measurements. (Experimental Error)
- True value - most probable value or accepted value based on references.

Answer the questions below based on your understanding of errors.

1. The freezing point of water is 273.2 K, but it was measured at 250.1 K. What is the percentage error?
2. The mass of a penny is 2.67 g, but it was measured at 2.55 g. What is the percentage error?
3. The air pressure was 101.3 kPa, but the weather forecaster said it was 10001.3 kPa. What is the percentage error?
4. The amount of heat released when 1 mole of CO₂ forms is 393.5 kJ, but it was measured at 378.2 kJ. What is the percentage error?
5. While doing a lab, a student found the density of a piece of pure aluminum to be 2.85 g/cm³. What was the student's percent error? (Hint: check reference table S.)