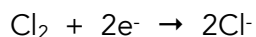
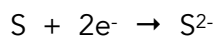


Unit 11 - Topic 2

Redox Half Reactions & Skeleton Reactions

Writing Half Reactions

Reduction: occurs when an atom or ion **gains** one or more electrons. The charge on the atom becomes **more negative**, or in other words, is REDUCED in value! The following half reactions are examples of reduction:



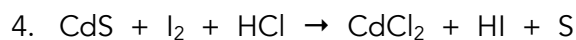
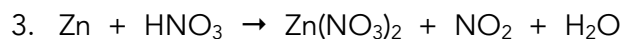
Oxidation: occurs when an atom or ion **loses** one or more electrons. The charge on the atom becomes **more positive**. The following half reactions are examples of oxidation:



Identify the following half reactions as oxidation or reduction. THEN, complete the reaction showing electrons in the right place (see examples above!).



Write the half reactions for each redox reaction below:



Name: _____

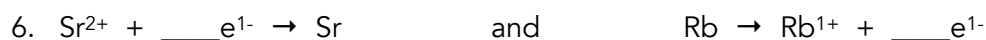
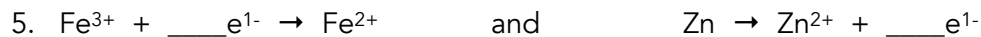
Date: _____

Balancing Skeleton Redox Reactions

Identify which is the oxidation half reaction and which is the reduction half reaction.

Explain or show how: # of electrons lost = # of electrons gained.

Combine the half reaction pairs in order to write a **balanced** skeleton reaction.



Write balanced half reactions for each of the following.

Indicate which half reaction is oxidation and which is reduction.

Combine the half reaction pairs in order to balance the given skeleton reactions.

