Date: _____



Unit 3 - Topic 5 Balancing Equations

1)
$$AlBr_3 + K_2(SO_4) \rightarrow KBr + Al_2(SO_4)_3$$

2)
$$\operatorname{FeCl}_3 + \operatorname{Na(OH)} \rightarrow \operatorname{Fe(OH)}_3 + \operatorname{NaCl}$$

3)
$$H_2(SO_4) + Na(NO_2) \rightarrow H(NO_2) + Na_2(SO_4)$$

4)
$$CH_4 + _O_2 \rightarrow _CO_2 + _H_2O$$

5)
$$C_3H_8 + \dots O_2 \rightarrow \dots CO_2 + \dots H_2O$$

6)
$$C_8H_{18} + \dots O_2 \rightarrow \dots CO_2 + \dots H_2O$$

Name: _____



Using the given word statement, determine the chemical equation. Include phases where you can. Make sure all of the formulas are written correctly, then balance the equation.

1. Potassium metal and chlorine gas combine to form potassium chloride.

2. Magnesium reacts with water to make magnesium hydroxide and hydrogen gas.

3. Zinc reacts with sodium fluoride to produce zinc fluoride and sodium.

4. Carbon reacts with oxygen gas to produce carbon monoxide.

5. When sodium metal reacts with iron (II) chloride, iron metal and sodium chloride are formed.

